

ELECTRICAL ENERGY STORAGE



100% Customisable

Unique Electricity Energy Storage Solutions

HydraRedox Iberia S.L. develops tailor-made storage solutions to address specific technical and economic requirements. HydraRedox systems are fully customisable in terms of power, energy, voltage and amperage.

5 kW	50 MW
1 hour	>24 Hours
100 A	> 2.000 A
12 V	1.000 V
	100 A

TECHNICAL SPECIFICATIONS				
	Efficiency			
Round Trip DC Efficiency	85 %			
☐ At Rated Load	>80 – 85%			
At Lower Loads	>85 – 90%			
Current Efficiency	Up to 95%			
Reaction time / UPS	340µs			
Depth of Discharge	100%			
	Life			
Number of Cycles	Unlimited			
Expected Life	~30 Years			
Self-discharge	Negligible			
	Safety			
Operating Temperature	-20° / +40°C			
Operating Pressure	Atmospheric			
	Dimensions			
Power (per 40ft ctnr)	Up to 150kW			
Energy (per 40ft ctnr)	Up to 1MWh			
Monitoring				
Visualization, recording and transmission for each cell: Voltage, Amperage, State-Of Charge, Temperature, Pressure				

0	HIGHLY RELIABLE
X	EASY MAINTENANCE
	AUTOMATIC CONTROL
	SAFE OPERATION
(3)	ENVIRONMENTALLY NEUTRAL
	MODULAR AND SCALABLE
•	TURN-KEY SOLUTIONS

Unique technical characteristics make it suitable for a wide range of applications:

Multiple Applications					
Wind	Solar PV	Utilities	Commercial Industrial Agricultural Micro-Grids Large-Scale		
	tible with wables	Grid-scale	Residential		

HydraRedox Iberia S.L. Gran Vía 36, 1° izq., 50005 ZARAGOZA contacto@hydraredoxiberia.com Tel +34 976 228896 www.hydraredox.com





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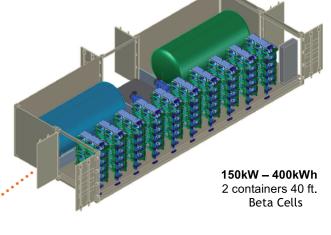
Revolutionary Vanadium Redox Technology

HydraRedox is an electric energy storage system based on a radically new approach to vanadium redox technology. The patented technology is based on an 'individual cell' design in which each cells operates independently and is constantly monitored and a unique electrolyte composition.

POWER SECTION (kW)

Individual electrochemical cells connected electrically in series convert chemical energy to electrical energy (and vice versa).



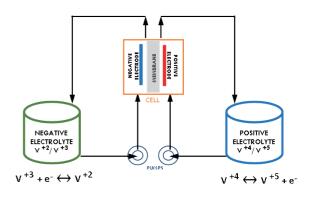


ENERGY SECTION (kWh)

The energy is stored in chemical form (the electrolytes) in two tanks (positive and negative) – and can be sized independently from power.

Vanadium Redox Technology Working Principles

Each cell is made of two compartments, one with a positive electrode, the other with a negative electrode, which are separated by a membrane. The electrodes come into contact with the electrolytes pumped from the tanks into the two compartments. Through a chemical reaction called **redox** (**reduction-oxidation**) the composition of the electrolytes (in terms of states of oxidation of **Vanadium**) changes, creating a shortage of electrons at the positive terminal (positive electrode) and a surplus at the negative terminal (negative electrode).



The charge and discharge reactions occur on the surface of the electrodes. During the **discharging cycle** (when the battery supplies energy), electrons flow from the negative to the positive terminal, generating an electrical current. During the **charging cycle** (when the battery is accumulating electricity from external sources), an electrical current applied to the terminals reverses the redox reaction and the electrons flow from the positive to the negative terminal.

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